

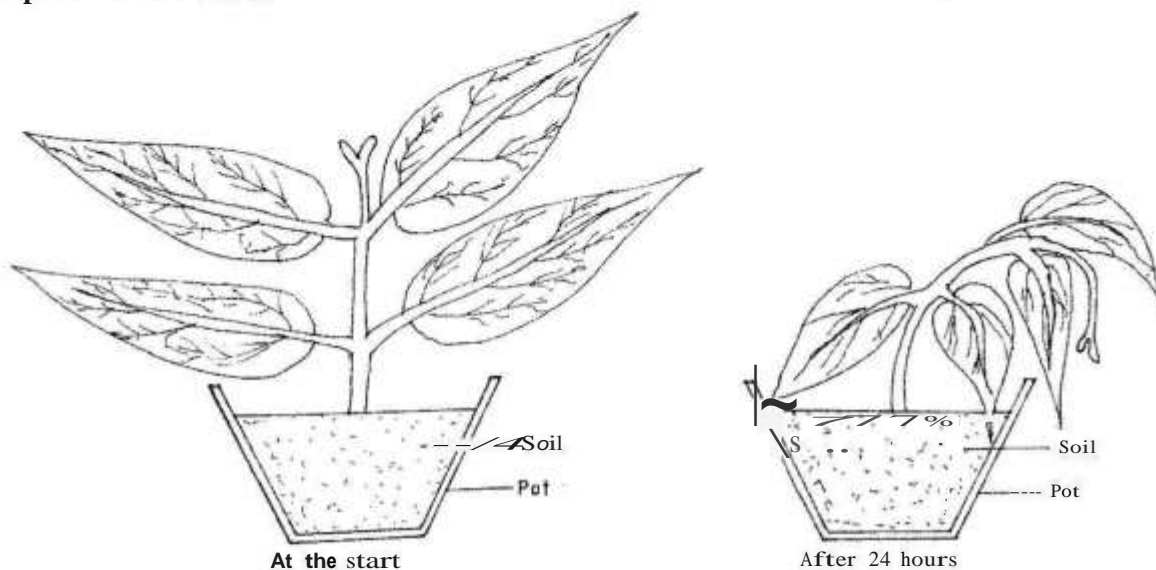
2.4 GENERAL SCIENCE (237)

2.4.1 General Science Paper 1 (237/1)

SECTION A: BIOLOGY (34 marks)

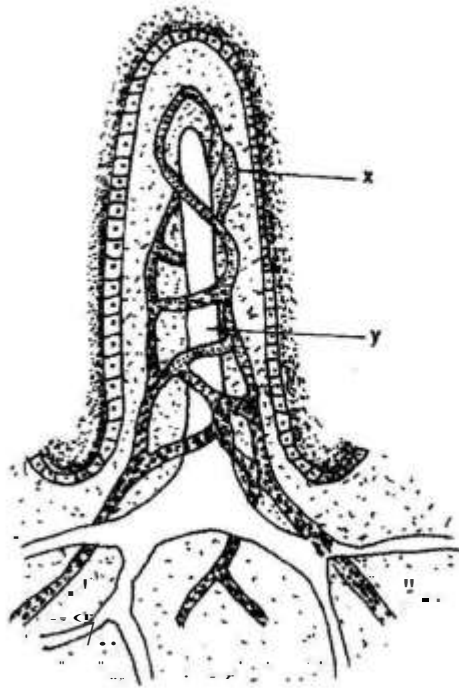
Answer *all* the questions in the spaces provided.

- 1 (a) Name the branch of biology that deals with the study of animals. (1 mark)
- (b) Give **two** reasons for classifying living organisms. (2 marks)
- (c) Give a reason why respiration is important in living organisms. (1 mark)
- 2 (a) Name the organelles observed under the light microscope in plant cells but not in animal cells. (2 marks)
- (b) State **two** precautions that should be taken when placing a microscope on a table for use. (2 marks)
- 3 In an experiment a solution was poured around a potted plant and left for 24 hours. The set up is as shown below.



- (a) State the nature of the solution. (1 mark)
-
- (b) Explain the observations made after the 24 hours. (2 marks)
- 4 (a) Name **one** raw material for photosynthesis. (1 mark)
- (b) **With** the exception of guard cells, how are the upper epidermal cells of a leaf adapted to their function? (2 marks)

5. (a) Distinguish between ingestion and egestion? (1 mark)
 (b) The diagram below represents a villus of the human alimentary canal.



Name the substances that are absorbed through the structures labelled **X** and **Y**.

(2 marks)

X:

Y:

- 6 Explain why a person in a poorly ventilated room with a burning charcoal stove may suffocate to death. (3marks)

- 7 (a) Name the causative agent of pneumonia. (1 mark)

.....

- (b) State **two** characteristics of an efficient respiratory surface which are absent in an amoeba. (2 marks)

- (c) How is anaerobic respiration applied in the baking industry? (1 mark)

- 8 (a) Explain how euphorbia leaves affect its transpiration. (2 marks)

- (b) Explain **one** adaptation of the root hair to its function. (2 marks)

- 9 (a) State **two** functions of the kidney. (2 marks)

- (b) Name **two** metabolic wastes removed through the skin. (2 marks)

- 10 State the effects of vasodilation. (2 marks)

SECTION B : CHEMISTRY (33 marks)

Answer **all** the questions in this section in the spaces provided.

- 11 What is meant by the following terms?
Valency. (1 mark)

.....
Electron affinity. (1 mark)

- 12 The information in the table below gives atomic sizes of elements P, Q and R that belong to the same group in the periodic table. Study it and answer the question that follows.
P, Q and R are not the actual symbols of the elements.

Elements	Atomic size (nm)
P	0.19
Q	0.23
R	0.15

Which element has the highest ionisation energy? Explain. (2 marks)

- 13 (a) Why is air considered a mixture and not a compound? (1 mark)
- (b) Give **one** similarity between rusting and combustion. (1 mark)
- (c) Which reagent will decompose to give oxygen gas in the presence of manganese (IV) oxide? (1 mark)
- (d) Carbon (IV) oxide does not support burning yet a piece of burning magnesium ribbon continues to burn in a gas jar full of carbon (IV) oxide. Explain. (1 mark)
- 14 (a) Using dots (•) and crosses (x) to represent electrons, draw the structure of sodium chloride. (Na = 11.0, Cl = 17.0). (2 marks)
- (b) Name the type of bond in diamond. (1 mark)

.....
15 Describe how a student would obtain sand from a mixture of sand and sugar. (2 marks)

- 16 (a) State the purpose of **pH** scale. (1 mark)

.....
(b) Hydrochloric acid is a strong acid. Explain the meaning of a strong acid. (1 mark)

(c) Dilute hydrochloric acid was reacted with solid calcium carbonate in a test tube. Write a balanced chemical equation for the reaction. (1 mark)

(d) Give **two** disadvantages of washing clothes in hard water using soapy detergents. (2 marks)

17 (a) Describe a test that can be used to show presence of water in a substance. (2 marks)

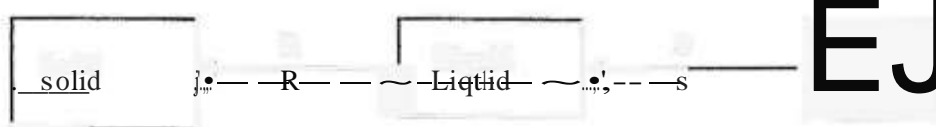
(b) Magnesium ribbon reacts slowly with water to form a solution that turns red litmus paper blue. Name the solution formed. (1 mark)

18 State how the following substances conduct electricity.

(a) Molten calcium chloride. (1 mark)

(b) Graphite. (1 mark)

19 The diagram below shows some changes in the physical states of matter. Study it and answer the questions that follows.



(a) Name the changes represented by letters R and S. (2 marks)

R:

S:

(b) Name the method used to separate coloured substances in a dye. (1 mark)

20 The grid below represents part of the periodic table. Use it to answer the questions that follow. The letters are not the actual symbols of the elements.

W										Z
	X								Y	

(a) Which letter represents an alkaline earth metal? (1 mark)

- (b) Which of the two elements represented by the letters W and Z is reactive. Explain. (2 marks)
- (iii) Write the electron arrangement of the element represented by Y. (1 mark)

- 21 (a) Name the type of reaction that occurs when a solution of lead (II) nitrate is added to a solution of sodium sulphate (in a boiling tube). (1 mark)

- (b) Write a balanced equation for the reaction that occurs when crystals of sodium nitrate are heated in a test tube. (1 mark)

- (c) Give the meaning of an acid salt. (1 mark)

SECTION C: PHYSICS (33 marks)

Answer *all* the questions in this section in the spaces provided.

- 22 **Figure 1**, shows a measuring instrument containing a liquid.

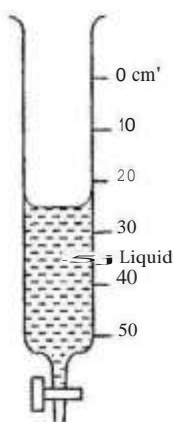


Figure 1

- (a) State the name of the instrument. (1 mark)
- (b) The reading changes from 26.8cm^3 to 42.5cm^3 after 40 drops of the liquid are released, determine the volume of each drop. (2marks)

- 23 (a) A block of wood is pulled along a horizontal surface. State **one** factor that determines the magnitude of the frictional force between the block and the surface. (1 mark)
- (b) **Figure 2** shows a vertical glass tube containing a liquid.

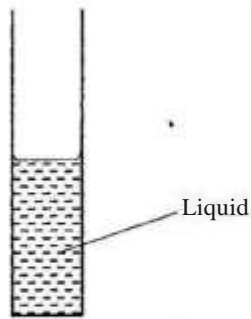


Figure 2

State a reason for the shape of the meniscus in terms of molecular forces. (2 marks)

- 24 (a) **Figure 3** shows two identical containers A and B filled with different liquids to the same level



Figure 3

It is observed that the pressure at the bottom of container B is higher than the pressure at the bottom of container A. State the reason for this observation. (1 mark)

- (b) Smoke particles inside a smoke cell are observed to move randomly when viewed through a microscope. Explain this observation. (2 marks)
- 25 A student observes that in the morning an overhead electrical power cable is straight and taut. At midday the student observes that the same cable has sagged. Explain these observations. (2 marks)

- 26 **Figure 4**, shows a crystal of potassium permanganate at the bottom of a beaker containing some water.

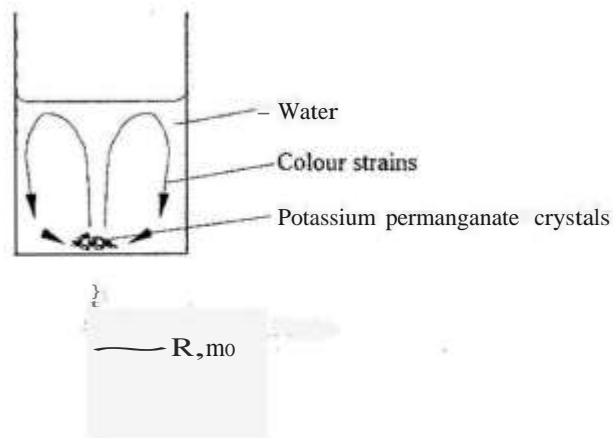


Figure 4

It is observed that when the beaker is heated from the bottom, strains of colour rise up from the crystal and curve out as shown. Explain this observation. (3 marks)

- 27 **Figure 5**, shows a uniform meter rule pivoted at the 60cm mark. The rule is balanced when a 40g mass is supported at the 90cm mark.

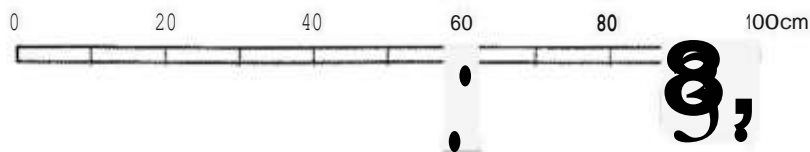


Figure 5

- (a) Show on the diagram the position of the centre of gravity of the metre rule. (1 mark)
 (b) Determine the mass of the metre rule. (2 marks)

- 28 **Figure 6**, shows a ball bearing resting on a flat surface.



Figure 6

- (a) Name the state of equilibrium of the ball bearing. (1 mark)

(b) State the reason for the answer in (a) above. (2 marks)

29 A spring extends by 20mm when supporting a mass of 150g. Determine the mass which when supported by the same spring causes an extension of 30mm. (3 marks)

30 A stone is thrown vertically upwards. Sketch a velocity - time graph for the motion of the stone from the time it is thrown until it comes back to the ground. (2 marks)

31 A box is lying on the floor of a fast moving lorry, it is observed that when the lorry is suddenly brought to rest, the final position of the box is nearer to the front of the lorry than before. Explain this observation. (2marks)

32 **Figure 7**, shows a pulley being used to raise a load.

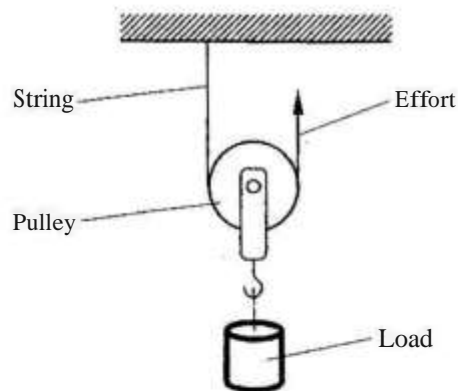


Figure 7

(a) State the velocity ratio of this machine. (1 mark)

(b) State **two** factors that reduce the efficiency of *this* machine. (2 marks)

33 (a) State the law of floatation. (1 mark)

(b) A ship is made of steel. Explain why the ship is able to float on water whose density is less than the density of steel. (2marks)