

**CHEMISTRY PAPER
FORM 1
TERM 3 2017
2017**

MARKING SCHEME

1. (a) To circle the 2 dots to the right.
(b) Line above – solvent front
Line below – base line
2.
 - Add water to the mixture, warm, stir
 - Filter the mixture, the residue is copper(II)oxide
 - Evaporate water and sodium chloride remains
3.
 - (a)
P – sublimation
Q – condensation
R – melting
S – deposition
 - (b)
 - Dry ice
 - Iodine
4.
 - (a) Increase, zinc react with air to form zinc oxide.
 - (b) Decrease – carbonate decompose to zinc oxide and carbon(IV)oxide which escapes into the air.
5.
 - i)
 - ii) Measure the temperature of the distillate coming out.
 - iii) C , it has the least boiling point, hence it boils faster.
6. To draw dropper.
7.
 - a) Sodium bromine
 - b) Zinc and sulphur
 - c) Lead and oxygen

Ecolebooks.com

- d) Magnesium and nitrogen
 - e) Potassium and iodine
- 8.

Element	Symbol
Carbon	C
Nitrogen	N
Oxygen	O
Hydrogen	H
Copper	Cu
Magnesium	Mg

- 9.
- Transparent, hence reactions can be observed easily.
 - Does not react with most reagents.
- 10.
- Does not make beakers dirty because it has no soot.
 - It is the hottest hence heats faster.
- 11.
- a) 30°C
- b)
- PQ – Temperature increases steadily as solid X absorbs heat energy, the heat absorbed increases kinetic energy of the particles as they vibrate vigorously.
- QR – solid-liquid Temperature remains constant until solid X melts heat supplied is used to weaken forces of attraction holding particle of solid X together. The particles move far apart as solid changes from solid to
- RS – Temperature rises steadily as the liquid X absorbs heat energy. The heat supplied increases further the kinetic energy of the particle causing them to move fast.
- 12.
- Bhang
 - Cocaine
 - Miraa
13. Wash immediately with cold water.
- 14.
- Don't run in the lab.
 - Label reagents to avoid confusion.
 - Don't smell gases directly, waft instead.
- 15.
- (a) Parts of the flame
- A – pale blue zone
- B – green blue zone
- C – almost colourless zone

(b)